Practice Problems: Atomic Mass Calculations I

See the complete solutions to these problems at https://www.youtube.com/watch?v=D0MEFWtoQDU.

1. Calculate the atomic mass of an element with two naturally occurring isotopes: ^{85}X (72.15%, 84.9118 amu) and ^{87}X (27.85%, 86.9092 amu). What is this element?

2. Chromium has the following isotopic masses and relative abundances. Determine the atomic mass of chromium to two decimal places.

Mass Number	Iostopic Mass (amu)	Percent Abundance
50	49.9461	4.35%
52	51.9405	83.79%
53	52.9407	9.50%
54	53.9389	2.36%